

Water and living organisms

1. The alkaline properties of solution with a pH >7 are due to:
 - hydroxide ions
 - hydroxonium ions
 - oxide ions
 - hydrogen ions
2. The acidic properties of solution with a pH <7 are due to:
 - hydroxide ions
 - oxide ions
 - hydroxonium ions
 - hydrogen ions
3. In a water molecule, each bond between the oxygen atom and the two hydrogen atoms are best described as:
 - a hydrogen bond
 - a single polar covalent bond, with a small positive charge on the hydrogens and a small negative charge on the oxygen
 - a single non-polar covalent bond, with a small negative charge on the hydrogens and a small positive charge on the oxygen
 - a single non-polar covalent bond